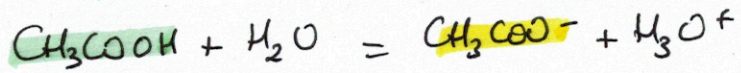


acide faible

base faible

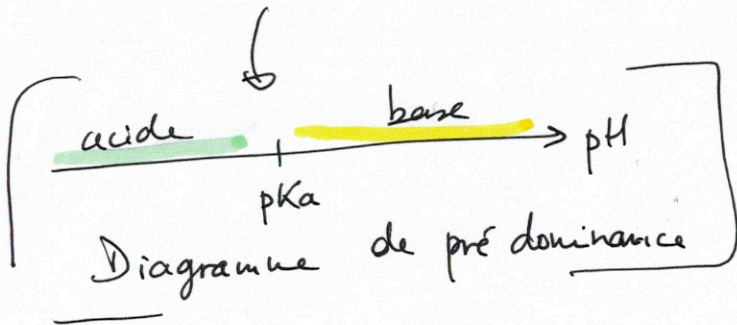


$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}_3\text{O}^+]}{[\text{CH}_3\text{COOH}]}$$
 constante d'acidité

$$\text{p}K_a = -\log K_a$$

$$\text{pH} = \text{p}K_a + \log \frac{[\text{base}]}{[\text{acide}]}$$

ACIDE - BASE



$$K = \frac{[\text{NH}_4^+][\text{HO}^-]}{[\text{NH}_3][\text{H}_3\text{O}^+]} = \frac{K_e}{K_a}$$

avec $K_e = [\text{H}_3\text{O}^+][\text{HO}^-] = 10^{-14}$ produit ionique de l'eau

